Solubility-Product Constants for Compounds at 25°C		
NAME	**FORMULA**	**K_{sp}**
Barium carbonate	$BaSO_4$	5.0×10^{-9}
Barium chromate	$BaCrO_4$	2.1×10^{-10}
Barium fluoride	BaF_2	1.7×10^{-6}
Barium oxalate	BaC_2O_4	1.6×10^{-6}
Barium sulfate	$BaSO_4$	1.1×10^{-10}
Cadmium carbonate	$CdCO_3$	1.8×10^{-14}
Cadmium hydroxide	$Cd(OH)_2$	2.5×10^{-14}
Cadmium sulfide*	CdS	8.0×10^{-28}
Calcium carbonate (calcite)	$CaCO_3$	4.5×10^{-9}
Calcium chromate	$CaCrO_4$	7.1×10^{-4}
Calcium fluoride	CaF_2	3.9×10^{-11}
Calcium hydroxide	$Ca(OH)_2$	6.5×10^{-6}
Calcium phosphate	$Ca_3(PO_4)_2$	2.0×10^{-29}
Calcium sulfate	$CaSO_4$	2.4×10^{-5}
Chromium (III) hydroxide	$Cr(OH)_3$	1.6×10^{-30}
Cobalt (II) carbonate	$CoCO_3$	1.0×10^{-10}
Cobalt (II) hydroxide	$Co(OH)_2$	1.3×10^{-15}
Cobalt (II) sulfide*	CoS	5.0×10^{-22}
Copper (I) bromide	$CuBr$	5.3×10^{-9}
Copper (II) carbonate	$CuCO_3$	2.3×10^{-10}
Copper (II) hydroxide	$Cu(OH)_2$	4.8×10^{-20}
Copper (II) sulfide*	CuS	6.0×10^{-37}
Iron (II) carbonate	$FeCO_3$	2.1×10^{-11}
Iron (II) hydroxide	$Fe(OH)_2$	7.9×10^{-16}
Lanthanum fluoride	LaF_3	2.0×10^{-19}
Lanthanum iodate	$La(IO_3)_3$	6.1×10^{-12}
Lead (II) carbonate	$PbCO_3$	7.4×10^{-14}
Lead (II) chloride	$PbCl_2$	1.7×10^{-5}
Lead (II) chromate	$PbCrO_4$	2.8×10^{-13}

Lead (II) fluoride	PbF_2	3.6×10^{-8}
Lead (II) sulfate	$PbSO_4$	6.3×10^{-7}
Lead (II) sulfide*	PbS	3.0×10^{-28}
Magnesium hydroxide	$Mg(OH)_2$	1.6×10^{-12}
Magnesium carbonate	$MgCO_3$	3.5×10^{-8}
Magnesium oxalate	MgC_2O_4	8.6×10^{-5}
Manganese (II) carbonate	$MnCO_3$	5.0×10^{-10}
Manganese (II) hydroxide	$Mn(OH)_2$	1.6×10^{-13}
Manganese (II) sulfide*	MnS	2.0×10^{-53}
Mercury (I) chloride	Hg_2Cl_2	1.2×10^{-18}
Mercury (I) iodide	Hg_2I_2	1.1×10^{-28}
Mercury (II) sulfide*	HgS	2.0×10^{-53}
Nickel (II) carbonate	$NiCO_3$	1.3×10^{-7}
Nickel (II) hydroxide	$Ni(OH)_2$	6.0×10^{-16}
Nickel (II) sulfide*	NiS	3.0×10^{-20}
Silver bromate	$AgBrO_3$	5.5×10^{-5}
Silver bromide	$AgBr$	5.0×10^{-13}
Silver carbonate	Ag_2CO_3	8.1×10^{-12}
Silver chloride	$AgCl$	1.8×10^{-10}
Silver chromate	Ag_2CrO_4	1.2×10^{-12}
Silver iodide	AgI	8.3×10^{-17}
Silver sulfate	Ag_2SO_4	1.5×10^{-5}
Silver sulfide*	Ag_2S	6.0×10^{-51}
Strontium carbonate	$SrCO_3$	9.3×10^{-10}
Tin (II) sulfide*	SnS	1.0×10^{-26}
Zinc carbonate	$ZnCO_3$	1.0×10^{-10}
Zinc hydroxide	$Zn(OH)_2$	3.0×10^{-16}
Zinc oxalate	ZnC_2O_4	2.7×10^{-8}
Zinc sulfide*	ZnS	2.0×10^{-25}

* For a solubility equilibrium of the type $MS_{(s)} + H_2O_{(l)} \leftrightarrows M^{2+}_{(aq)} + HS^-_{(aq)} + OH^-_{(aq)}$